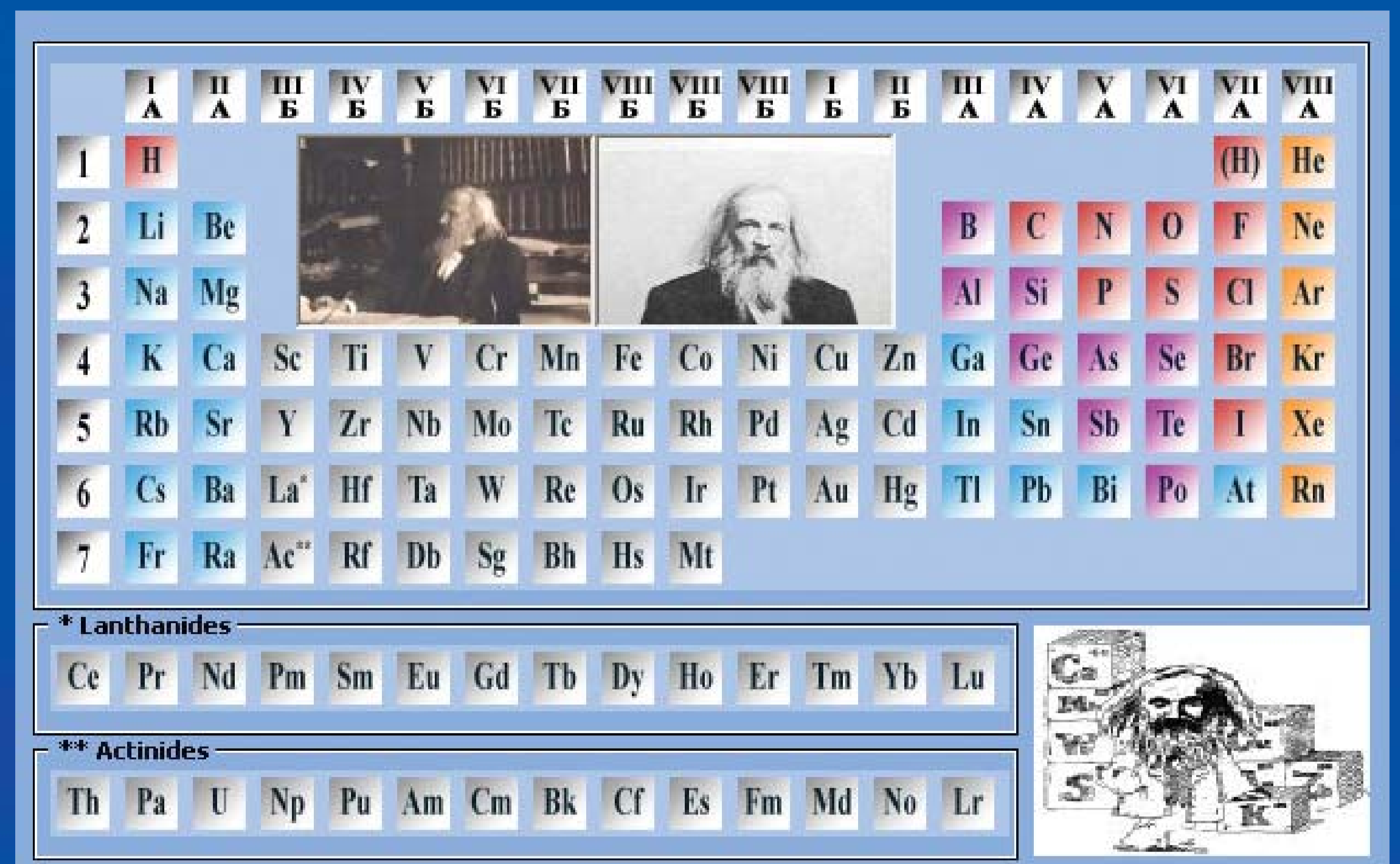
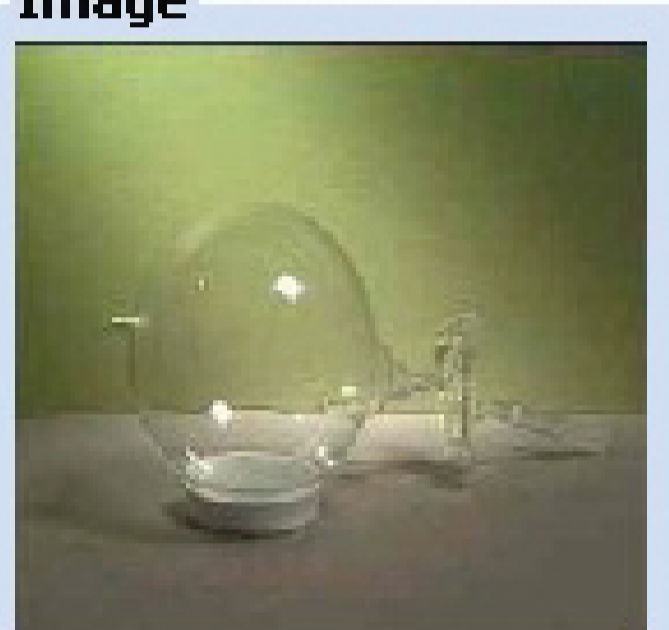


# Applications which can help you at any time. You can rely on them whatever you do.

The Periodic Table is a chart that lists elements by atomic number, so that elements with similar chemical properties are in the same column. Metals, non-metals, elements from B groups and inert gases are tinted in different colors.



**Image**



**Physical Properties**

Chemical sign: **O**  
 Number: 8  
 Atomic weight: 15,9994  
 Temperature of melting: -218.4  
 Temperature of boiling: -182,962  
 Electrons by layers: 2:6  
 Isotopes: 8  
 Atomic radius: 0,65

**History**

Oxygen is a Group 16 element. While about one fifth of the atmosphere is oxygen gas, the atmosphere of Mars contains only about 0.15% oxygen. Oxygen is the third most abundant element found in the sun, and it plays a part in the carbon-nitrogen cycle, one process responsible for stellar energy production. Oxygen in excited states is responsible for the bright red and yellow-green colours of the aurora. About two thirds of the human body, and nine

ors.

When

you click on an element you get detailed information about it: atomic weight, melting and/or boiling temperature, atomic radius, electrons by layers, etc. There is also additional interesting information about the element and a picture which shows either the element or its usage.

The Solubility Table lets you check whether two ions (a cation and an anion) can form a substance and whether it would be soluble or insoluble.

Ions	Names	H <sup>+</sup>	NH <sub>4</sub> <sup>+</sup>	K <sup>+</sup>	Na <sup>+</sup>	Ag <sup>+</sup>	Ba <sup>2+</sup>	Ca <sup>2+</sup>	Mg <sup>2+</sup>	Zn <sup>2+</sup>	Cu <sup>2+</sup>	Hg <sup>2+</sup>	Pb <sup>2+</sup>	Fe <sup>2+</sup>	Fe <sup>3+</sup>	Al <sup>3+</sup>
OH <sup>-</sup>	hydroxide	P	P	P	P	-	P	M	H	H	H	-	H	H	H	H
NO <sub>3</sub> <sup>-</sup>	nitrate	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P
S <sup>2-</sup>	sulphide	P	P	P	P	H	-	-	-	H	H	H	H	H	H	H
SO <sub>3</sub> <sup>2-</sup>	sulphite	P	P	P	P	M	M	M	M	-	-	-	H	M	-	-
SO <sub>4</sub> <sup>2-</sup>	sulphate	P	P	P	P	M	H	M	P	P	P	P	H	P	P	P
CO <sub>3</sub> <sup>2-</sup>	carbonate	P	P	P	P	H	H	H	H	H	H	H	H	H	-	-
SiO <sub>3</sub> <sup>2-</sup>	siliconate	H	-	P	P	H	H	H	H	H	H	H	H	H	H	H
PO <sub>4</sub> <sup>3-</sup>	phosphate	P	P	P	P	H	H	H	H	H	H	H	H	H	H	H
F <sup>-</sup>	fluoride	P	P	P	P	P	M	H	H	M	H	M	H	M	P	P
Cl <sup>-</sup>	chloride	P	P	P	P	H	P	P	P	P	P	P	M	P	P	P
Br <sup>-</sup>	bromide	P	P	P	P	H	P	P	P	P	P	P	M	P	P	P
I <sup>-</sup>	iodide	P	P	P	P	H	P	P	P	P	P	H	H	P	P	P
CrO <sub>4</sub> <sup>2-</sup>	chromate	P	P	P	P	H	H	M	P	P	H	H	H	P	P	M
CH <sub>3</sub> COO <sup>-</sup>	acetate	P	P	P	P	P	P	P	P	P	P	P	P	P	P	P

P Soluble   
 M Less soluble   
 H Insoluble   
 - Breaking down   
 S Sediment

able. If the two ions form a sediment you can see its color.

**Search**

Enter a word:

**List with all words and terms**

- Absolute Error
- Absolute Zero
- Accuracy
- Acid
- Actinides
- Activation Energy
- Alkali Earth Metal
- Alkali Metal**
- Alpha Particle
- Anion
- Atom
- Atomic Mass
- Base
- Beta Particle
- Boiling Point
- Cation
- Ceiling

**[\*] ALKALI METAL**

A metal in the first column of the periodic table (i.e., lithium, sodium, potassium, rubidium, cesium and francium). With the exception of francium, these metals are all soft and silvery. They may be readily fused and volatilized with their melting and boiling points becoming lower with increasing atomic mass. They are the strongest electropositive metals. These elements react vigorously, even violently with water.

When you are not used to chemical terminology you find it hard to remember without constantly using it. This will no longer be a problem as the program has a built-in glossary of the most widely used terms. It can easily refresh your memory about the meaning of an unfamiliar term at any time.